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### [CSE543/FALL 2006 - QUIZ](#)

Thursday, November 16, 2006 Professor Trent Jaeger Please read the instructions and questions carefully. You will be graded for clarity and correctness. You have 25 minutes to complete this quiz, so focus on those questions whose subject matter you know well. Write legibly and check your answers before handing it in. Short Answer - some will be one or two words no more than 3 sentences 1. (3pts) Why ...

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### [MICROSOFT WORD - QUIZ 3 ANSWERS.DOCX](#)

Chemistry 31 Quiz 3 10 minutes March 20 Name Lab Section Number H2C2O4  $K_{a1} = 5.62 \times 10^{-2}$ ,  $K_{a2} = 5.42 \times 10^{-5}$ ;  $K_w = 1.00 \times 10^{-14}$ . (5 pts) Calculate the pH of the following (Ignore activity and assume no secondary reactions besides acid base in water.): 0.015 M Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> [C<sub>2</sub>O<sub>4</sub><sup>2-</sup>] = 0.015 M See Chapter 6 problems 46, 48, and 6-1  $1.00 \times 10^{-14}$  C<sub>2</sub>O<sub>4</sub><sup>2-</sup> + H<sub>2</sub>O ? HC<sub>2</sub>O<sub>4</sub><sup>-</sup> + OH<sup>-</sup>  $K_{b1} = 5 = 1.845 \times 10^{-10}$   $5.42 \times 10^{-5}$  [HC<sub>2</sub>O<sub>4</sub><sup>-</sup>] [OH<sup>-</sup>]

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